Oxidation Numbers & Balancing Redox Equations

- 1. c
- 2. b
- 3. c
- 4. b
- 5. c
- 6. a
- 7. b
- 8. b
- 9. b
- 10.

a)
$$5 \text{ CO} + I_2O_5 \rightarrow 5 \text{ CO}_2 + I_2$$

b)
$$2 \text{ H}_2\text{O} + \text{NiO}_2 + 2 \text{ e}^- \rightarrow \text{Ni(OH)}_2 + 2 \text{ OH}^-$$

$$2 \text{ OH}^- + \text{Fe} \longrightarrow \text{Fe}(\text{OH})_2 + 2 \text{ e}^-$$

c)
$$CO_2 + H_2O + 2 e^- \rightarrow CO + 2 OH^-$$

$$2 \text{ OH}$$
 + $2 \text{ NH}_2\text{OH} \longrightarrow \text{N}_2 + 2 \text{ e}$ + $4 \text{ H}_2\text{O}$